

Vocabulary/Study Guide - Periodic Table

The test is all multiple choice and has 60 to 70 questions.

It must be completed in ONE CLASS PERIOD.

You may NOT use your notes or a calculator.

Items in **BLUE** are not directly asked about on the test but are helpful to know.

A groups / B groups

actinide series

alkali metal

alkaline earth metal

anion

atomic radius

cation

characteristics of metals vs non-metals

Color coding on periodic table

Dividing line between metals and nonmetals

Dot notation related to position in the table

electron affinity

electronegativity

element 104 - Rf - rutherfordium, created 1969 (Named after Ernest Rutherford)

element 105 - Ha - hahnium, created 1970 (Named after Otto Hahn)

element 106 - Sg - seaborgium, created 1974 (Named after Glenn Seaborg) * Officially Named in 1994

element 107 - Ns - nielsborhium, created 1976 (Named after Niels Bohr)

element 108 - Hs - hassium, created

element 109 - Mt - meitnerium, created

family/group

halogen

inert

inner-transition element

ionization energy (1st, 2nd, 3rd, etc...)

Labs: Periodic Properties-Cards; Oxygen gas Lab; Alkaline Element Lab; Graphing Periodic Trends

lanthanide series

Mendeleev

metal / non-metal / metalloid (semi-metal)

metallic vs non-metallic properties (conductivity, ionization energy, etc...)

Moseley

most reactive family of metals

most reactive family of non-metals

noble gas

noble gas configuration

octet

octet rule

period

periodic law

periodic trend / group trend

periodicity

Ramsay, William - Discoverer of 4 Noble gases

representative elements

s,p,d,f blocks

shielding

transition element

transuranic element

trends of: atomic radius, ionization energy, metallic character, etc...(and explain reasons for)

triads, telluric helix & 'Law of Octaves' (and who noticed them)

valence electron

zero group (noble gases or group 18)